Disclosed are carbenes of the general formula:

and including salts thereof, and metal complexes thereof. The carbenes are useful in any reaction where carbenes and carbone-metal complexes are used. The carbenes disclosed herein are particularly useful in asymmetric catalysis.

33 Claims, 10 Drawing Sheets
OTHER PUBLICATIONS

Roettz et al., (2003), Binaphthyldiamine-Based Diazaphospholidines as a New Class of Chiral Monodentate P-Ligands, Synthesis, 12:1809-1814.

STN search (Nov. 24, 2009).


FIG. 1
FIG. 10
This is a continuation of application Ser. No. 11/102,964, filed Apr. 11, 2005, now U.S. Pat. No. 7,750,149, issued Jul. 6, 2010.

BACKGROUND


A variety of heterocyclic frameworks have been employed in the preparation of NHCs, including four-, five- and six-membered rings, as represented by compounds 1 through 5. To date, most catalytic applications employ the five-membered analogs.

There remains, however, a clear and unmet need for chiral NHCs that can be used as catalysts or ligands in asymmetric synthesis. The ability to synthesize, isolate, and prepare metal complexes of NHCs that possess a non-planar heterocyclic framework has significant utility toward this end.

SUMMARY OF THE INVENTION

The present invention is directed to 7-membered, N-heterocyclic carbenes, salts thereof, and metal complexes thereof. The carbenes are generally of formula 1, metal complexes of formula 1, and salts of formula 1:
the carbon to which they are attached and independent of R.

Ring A and ring B are independently selected from the group consisting of substituted or unsubstituted mono- or polycyclic cycloalkenyl, cycloaryl, heterocycloalkenyl, and heterocycloalkynyl having up to 34 atoms within each of ring A or ring B, wherein heteroatoms are in either of ring A or ring B, if any, are independently selected from the group consisting of N, O, S, and P.

R¹, R², R³, R⁴, R⁴, and R⁶ are independently selected from the group consisting of hydrogen; C₁ to C₆₀ substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl; aryl and heteroaryl; or R¹, R², and R³, including or excluding the carbon to which they are attached and independent of R⁴, R⁵, and R⁶, including or excluding the carbon to which they are attached and independent of R⁴, R⁵, and R⁶, define a C₃ to C₆₀ , substituted or unsubstituted, mono- or polycyclic cycloalkenyl, cycloalkynyl, cycloaryl, heterocycloalkenyl, heterocycloalkynyl, wherein heteroatoms in, if any, are independently selected from the group consisting of N, O, S, and P.

Where substituents are present on any of rings A and B or any of the R moieties, the substituents are selected from the group consisting of halogen; linear or branched C₁ to C₁₂ -alkyl, cycloalkyl, cycloalkenyl, or cycloalkynyl; mono- or polycyclic aryl, and mono- or polycyclic heteroaryl having up to 34 atoms within each of ring A or ring B, wherein heteroatoms, if present, are independently selected from the group consisting of N, O, S, and P.

The present invention explicitly encompasses the free carbenes (as illustrated in formula I), metal complexes of formula I, and salts of formula I.

In the preferred embodiment, ring A and ring B are selected from the group consisting of phenyl, naphthyl, anthracenyl, phenanthrenyl, and pyrenyl, pyridyl, pyrroldinyl, quinoline, indole, and thiophene. Also in the preferred embodiment, R¹, R², R³, R⁴, R⁵, and R⁶ are independently selected from the group consisting of hydrogen; C₁ to C₆₀ substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl.

The present invention also encompasses carbenes formula II, metal complexes of formula II, and salts of formula II:

FORMULA II

Here, R¹ through R⁶ are as defined previously.

R⁷, R⁸, R⁹, R₁₀, R₁₁, R₁₂, R₁₃, and R₁₄ are independently selected from the group consisting of hydrogen, halide, C₁ to C₆₀ substituted or unsubstituted, linear or branched alkyl, alkynyl, and alkynyl, aryl, heteroaryl, cyano, thiolate, alkoxy, primary amido, and secondary amido; or any of R¹, R², R³, and R⁴ combined and any of R¹¹, R₁₂, R₁₃, and R₁₄ combined independently define substituted or unsubstituted mono- or polycyclic cycloalkenyl, cycloaryl, heterocycloalkenyl, heterocycloalkynyl fused rings having up to 34 atoms; wherein heteroatoms, if present, are independently selected from the group consisting of N, O, S, and P.

More specifically still, the carbenes of the present invention comprise formula III complexes:

FORMULA III

Here, ring A, ring B, R¹, R², R³, R⁴, R⁵, and R⁶ are as defined as defined previously. M is hydrogen or a metal. L (when present) is one or more ligand(s) coordinated to M. When there is more than one ligand (L), each L substituent may be the same or different [e.g., L₃ explicitly encompasses (L¹)(L²)(L³)]. X is any counter-ion (or group of counterions) without limitation, and “a” is an integer.

In the preferred formula III carbenes, M is hydrogen or a metal selected from the group consisting of Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Os, Ir, Pt, and Au, and X is selected from the group consisting of tetrafluoroborate, hexafluorophosphate, and halide.

In formulae I and III, the most preferred compounds are those wherein rings A and B are phenyl or naphthyl. Likewise, in formula II, the most preferred compounds are those wherein R¹, R², R³, R⁴, R⁵, and R⁶ are hydrogen, or wherein are hydrogen any of R¹, R², R³, and R⁴ combined, and any of R¹¹, R₁₂, R₁₃, and R₁₄ combined define a fused phenyl ring.

BRIEF DESCRIPTION OF THE FIGURES

FIG. 1 is a depiction of the molecular structure of compound 10 as revealed by single-crystal X-ray diffraction. The hydrogen atoms and the BF₄-counterion are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

FIG. 2 is a depiction of the molecular structure of compound 11 as revealed by single-crystal X-ray diffraction. The hydrogen atoms are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

FIG. 3 is a depiction of the molecular structure of the palladium complex 14 as revealed by single-crystal X-ray diffraction. The hydrogen atoms and solvent molecules are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

FIG. 4 is another depiction of the molecular structure of the palladium complex 14 as revealed by single-crystal X-ray diffraction. The hydrogen atoms are omitted for clarity. The allyl ligand is disordered in a 73:27 ratio.

FIG. 5 is a depiction of the molecular structure of the palladium complex 15 as revealed by single-crystal X-ray diffraction. The hydrogen atoms and solvent molecules are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

FIG. 6 is a depiction of the molecular structure of the asymmetric unit of the palladium complex 15 as revealed by
single-crystal X-ray diffraction. The hydrogen atoms are omitted for clarity. Thermal ellipsoids are shown at 30% probability.

FIG. 7 is a depiction of the molecular structure of complex 16 as revealed by single-crystal X-ray diffraction. The hydrogen atoms except for those on the water ligand are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

FIG. 8 is a depiction of the molecular structure of complex 17 as revealed by single-crystal X-ray diffraction. The hydrogen atoms except for those on the water ligand are omitted for clarity. Thermal ellipsoids are shown at 50% probability. Only the preferred orientation is shown.

FIG. 9 is another depiction of molecular structure of the (NHC)-Pd(O₂CF₃)(OH₂) complex 17 as revealed by single-crystal X-ray diffraction. The hydrogen atoms except for those on the water ligand are omitted for clarity. Thermal ellipsoids are shown at 50% probability. The NHC ligand is disordered.

FIG. 10 is a depiction of the molecular structure of the (NHC)-Ag-OAc complex 18 as revealed by single-crystal X-ray diffraction. The hydrogen atoms are omitted for clarity. Thermal ellipsoids are shown at 50% probability.

**DETAILED DESCRIPTION OF THE INVENTION**

Definitions

The following definitions are provided for sake of clarity. Terms not explicitly defined are to be given their accepted definitions in the field of organic chemistry in general, and the field of synthetic organic chemistry in particular.

Where compounds are designated by chemical structure drawings, without any designation of stereoisomerism or positional isomerism (e.g., cis vs. trans isomerism), the structure explicitly encompasses all such stereoisomers, positional isomers, enantiomers, diastereomers, enantiomerically pure or enantiomerically enriched forms thereof, and racemic mixtures thereof.

The term “metal” refers to any metal on the periodic chart of the elements. Transition and main group metals of Group 3 to Group 13 are preferred, and metals of Groups 8-11 (Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Os, Ir, Pt and Au) are most preferred.

Regarding salts of the subject compounds, the subject compounds disclosed herein are carbenes. As such, the carbene carbon is a neutral species, having two single bonds to neighboring atoms (nitrogen atoms in this instance) and a pair of non-bonding electrons. In its neutral state, however, the carbene carbon has only six electrons, and is stabilized by donation of adjacent nitrogen lone pair electron density to the empty carbon orbital to alleviate the electron deficiency. Coordination of a proton to the nonbonding pair of electrons of the carbene results in the formation of an amidinium salt that may possess any of numerous possible counter-anions, e.g. halides, tetrafluoroborate, hexafluorophosphate, etc.

Chemistry:

Exemplary of the subject compounds is the axially chiral, seven-membered N-heterocyclic carbene (NHC), compound 8:

where R is adamantyl.

The initial synthetic efforts targeted analogs of 8 with bulky aryl substituents (e.g., R=mesityl; 2,6-diisopropylphenyl), which are commonly found in stable NHCs. Compound 9H was obtained in quantitative yield by reduction of 2,2'-dinitrobiphenyl, and mesitylation of 9H proceeds effectively under previously reported conditions. See Reetz et al. (2003) *Synthesis* 12:1809-1814. The subsequent condensation step to form the amidinium salt, however, was unsuccessful using a variety of known protocols:

The origin of this failure remains uncertain, and while not being limited to any particular mechanism, it is thought that the weak basicity of the diaryl amines may cause kinetic or thermodynamic problems in this reaction. Therefore, N-alkylated derivatives were proposed as potential alternatives.

Preparation of 9H with sterically encumbered primary and secondary alkyl substituents proceeds smoothly (R=neopentyl, 2-adamantyl). The condensation of 9H with 2-adamantanone, followed by reduction of the diimine with lithium aluminum hydride produces 9₅⁻::₄ in quantitative yield. (See the Examples for full experimental details.) Heating this compound in neat triethylorthoformate with NH₄BF₄ (Saba et al. (2001) *J. Chem. Soc., Perkin Trans. 1*, 1586-1593)
produces the racemic C$_2$-symmetric amidinium tetrafluoroborate salt, 10, which was characterized by single crystal X-ray diffraction (see FIG. 1 and the Examples).

11 was prepared from 9$^\beta$, in a related manner. Pivaloyl chloride was added to 9$^\beta$, and subsequent reduction of the carbonyl group by lithium aluminum hydride gave 9$^\beta$ in 92% overall yield. Amidinium salt 11 was prepared in the same manner as 10 in 93% yield (see FIG. 2 and the Examples). In the same way, the non-racemic salt (R)-13 was synthesized from (R)-12$^\beta$ in three steps, with an overall yield of 42% (see the Examples).

The amidinium salts 10, 11, and (R)-13 represent attractive NHC precursors. Efforts were then focused on preparing NHC-coordinated palladium(II) complexes. Deprotonation of 10 in a THF solution of [Pd(allyl)Cl]$_2$ (Jensen & Sigman (2003) Org. Lett. 5:63-65, Viciu et al. (2004) Organometallics 23:1629-1635) generates the air-stable NHC–Pd(allyl)Cl complex, 14, in high yield. The $^1$H NMR spectrum of 14 reveals the presence of two diastereomeric allyl rotamers in...
solution in a 1.4:1 ratio. Single-crystal X-ray diffraction studies confirmed the structure of 14 (see FIG. 3) and revealed that a mixture of isomers is also present in the solid state, with the allyl group disordered over two positions in a 2.7:1 ratio. See FIG. 4 and the Examples.

Protonolysis of the allyl ligand of 14 with HCl in diethyl ether produces [NHC-Pd(Cl)]$_2$, 15, in quantitative yield. Only one isomer of this dimeric compound is detected in solution by $^1$H NMR spectroscopy. Single crystal X-ray diffraction studies reveal the presence of a heterochiral dimer (see FIGS. 5 and 6 and the Examples). Treatment of 15 with silver acetate or silver trifluoroacetate in wet CH$_2$Cl$_2$ yields the NHC—Pd carboxylate complexes 16 and 17, respectively (see FIGS. 7, 8 and 9).

Synthesis of NHC-Coordinated Pd(II) Complexes
Reaction of 11 with potassium tert-butoxide in the presence of Pd(OAc)_{2} and NaI, followed by addition of AgOAc resulted in the formation of NHC—Ag—OAc complex 18 in 18% yield (See FIG. 10 and the Examples).

Rzepa and coworkers previously reported a purely computational analysis of 7-membered, 8π-electron NHCs. See Kastrup, Oldfield, & Rzepa, (2002) *J. Chem. Soc., Dalton Trans.* 2421-2422. These ab initio computational studies revealed that significant reduction of the expected antiaromatic character arises from twisting of the strained heterocyclic ring allowing for Möbius-type aromatic stabilization. See Hall & Rzepa (2003) *Org. Biomol. Chem.* 1:182-185; Heilbrunner (1964) *Tetrahedron Lett.* 29:1923-1928; and Zimmerman (1971) *Acc. Chem. Res.* 4:272-280. Rzepa’s investigations, however, were purely computational, and Rzepa’s group made no attempt to synthesize the compounds described therein. The compounds disclosed herein are the first metal complexes of a seven-membered NHC. They possess a large torsional twist resulting in axial chirality and are synthesized via a route amenable to the preparation of diverse analogs. The C_{6}-symmetric architecture can be readily modified with different nitrogen substituents, and the biaryl backbone can be altered to include rotationally hindered and enantiomerically pure derivatives, for example, binaphthyl.

Utility:

The carbenes described herein are useful in any protocol or reaction scheme for synthesizing chemical compounds wherein other NHCs are conventionally used, without limitation. In particular, NHCs find use as nucleophilic catalysts and as ligands coordinated to metals in homogeneous metal-catalyzed transformations.

For examples of NHCs used as nucleophilic catalysts, see Teles et al. (1996) *Helv. Chim. Acta* 79:61-83, in which NHCs are used to catalyze the condensation of two aldehydes to form β-hydroxyketones. The above reaction type has been carried out—employing chiral NHCs to give optically active β-hydroxyketones in good yields and enantioselectivities (See Enders et al. (1996) *Helv. Chim. Acta* 79:1217-1221).


The compounds disclosed herein can be used in any of the above-noted types of reactions.

The most preferred use for the present compounds, however, is as enantioselective catalysts. Metal complexes according to the present invention are highly useful as catalysts for alkene metathesis reactions. For analogous reactions with other types of carbenes, see Trnka & Grubbs (2001) *Acc. Chem. Res.* 34:18-29. Thus, for example, metal complexes of the present compounds can be used to catalyze the asymmetric hydrogenation of alkenes, the hydroisilylation of methyl ketones, and the hydroisilylation of acetylene complexes. Complexes according to the present invention can also be used to catalyze enantioselective ring-opening and ring-closing metathesis reactions. See, for example, Seiders et al. (2001) *Org. Lett.* 3:3225-3228, and Van Veldhuizen et al. (2002) *J. Am. Chem. Soc.* 124:4954-4955.

**EXAMPLES**

The following Examples are included solely to provide a more complete description of the invention disclosed and claimed herein. The Examples do not limit the scope of the claimed invention in any fashion.

General: All manipulations were performed under an inert nitrogen atmosphere unless otherwise specified. Dry, oxygen-free solvents were employed. 1H and 13C NMR spectra were recorded on either a Bruker Homer-300, a Bruker Athena-300, a Varian Mercury-300 or a Varian Inova-500 NMR spectrometer. 1H chemical shifts are reported in ppm relative to Me_{4}Si as an external standard, while 13C chemical shifts are reported in ppm relative to CHCl_{3}.

Synthesis of 2,2'-Diaminobiphenyl, Compound 9H: Compound 9H was synthesized by an adaptation of a literature procedure. Gillespie et al. (2002) *J. Org. Chem.* 67:3450-3458. 2,2'-Dinitrobiphenyl (102.0 g, 417.5 mmol) and 10% Pd/C (16.4 g) were combined with 300 mL EtOAc in a hydrogenation vessel. The vessel was pressurized to 40 psi H_{2} for 3.5 h (when H_{2} was no longer being consumed). The slurry was filtered through a plug of celite. Rotary evaporation followed by drying on a vacuum line gave pure product as light yellow/orange powder in 100% yield. 1H NMR (CDCl_{3}, 297K, 300 MHz): δ 3.71 (s, 4H), δ 6.79 (s, 4H), δ 6.84 (td, 2H, J=7.4, 1.2 Hz), δ 7.14 (dd, 2H, J=7.5, 1.2 Hz), δ 7.19 (td, 2H, J=7.5, 1.5 Hz). 13C NMR (CDCl_{3}, 297K, 300 MHz): δ 115.58, δ 118.71, δ 124.62, δ 128.81, δ 131.08, δ 144.22. HRMS (ESI-EMM): m/z=185.1073 ([M+H]^{+}), Δ=2.7 ppm.
Synthesis of Amidinium Salt, Compound (±)10: Compound (±)10 was synthesized according to a modified literature procedure, see Alder et al. (2001) J. Chem. Soc., Perkin Trans. 1, 1586-1593, and Saba et al. (1991) Tetrahedron Lett. 32:5031-4. 5.76 g (12.7 mmol) of 2,2'-bis(2-adamantylamino)biphenyl 9 and 1.3 g (12.7 mmol) NH$_4$BF$_4$ were combined under nitrogen in a 500 mL round-bottomed flask, and approx. 200 mL triethyl orthoformate was added. The reaction was heated to 100° C. for 16 h, after which time the product had crashed out of solution as a white powder. The product had 65% yield as a light fluffy white powder without further purification. Crystals suitable for X-ray analysis were achieved by vapor diffusion of n-pentane onto a CHC$_2$Cl$_2$ solution of (±)10. 1 H NMR (CDCl$_3$, 297K, 300 MHz): δ 0.97 (d, 2H, J=13.0 Hz), δ 1.14 (d, 2H, J=13.0 Hz), δ 1.47-2.09 (m, 22H), δ 2.66 (s, 2H), δ 4.77 (s, 2H), δ 7.43 (m, 2H), δ 7.51 (m, 6H), δ 8.84 (s, 1H). 13 C NMR (CDCl$_3$, 297K, 300 MHz): δ 26.39, δ 26.74, δ 30.00, δ 30.11, δ 30.36, δ 31.3, δ 31.7, δ 31.8, δ 32.2, δ 37.4, δ 37.8, δ 37.9, δ 56.7, δ 111.1, δ 116.3, δ 124.0, δ 129.1, δ 130.9, δ 145.3. HRMS (ESI-EMM): m/z=453.3247 ([M+BF$_4$]$^+), Δ=3.0 ppm.

Synthesis of 2,2'-Bis(2-Adamantylamino)Biphenyl, Compound 9: 614 mg (3.33 mmol) of 2,2'-diaminobiphenyl was combined with 35% (per amine functionality) pTsOH in a Dean-Stark apparatus. The reagents were dissolved in approximately 200 mL toluene and refluxed for 72 hours. The solvent was stripped off, and 122 mg (3.22 mmol) LAH was added, followed by approximately 200 mL THF. The reaction flask was heated to 50° C. for 2 h, followed by a careful quenching with approximately 100 mL water and 10 mL sat. NH$_4$Cl. The resulting slurry was filtered through a plug of celite, and the plug washed with CH$_2$Cl$_2$. The aqueous layer was washed once with CH$_2$Cl$_2$ (approx. 100 mL). The organic layers were combined, dried over MgSO$_4$, filtered, and the solvent was removed, yielding pure 9 in 100% yield. 1 H NMR (CDCl$_3$, 297K, 300 MHz): δ 1.38-1.97 (m, 28H), δ 3.54 (s, 2H), δ 4.05 (s, 2H), δ 6.71 (m, 4H), δ 7.11 (m, 2H), δ 7.22 (m, 2H). 13 C NMR (CDCl$_3$, 297K, 300 MHz): δ 27.4, δ 27.6, δ 31.3, δ 31.7, δ 31.8, δ 32.2, δ 37.4, δ 37.8, δ 37.9, δ 56.7, δ 111.1, δ 116.3, δ 124.0, δ 129.1, δ 130.9, δ 145.3. HRMS (ESI-EMM): m/z=453.3247 ([M+H]$^+$), Δ=5 ppm.

Synthesis of Amidinium Salt, Compound (±)10: Compound (±)10 was synthesized according to a modified literature procedure, see Alder et al. (2001) J. Chem. Soc., Perkin Trans. 1, 1586-1593, and Saba et al. (1991) Tetrahedron Lett. 32:5031-4. 5.76 g (12.7 mmol) of 2,2'-bis(2-adamantylamino)biphenyl 9 and 1.3 g (12.7 mmol) NH$_4$BF$_4$ were combined under nitrogen in a 500 mL round-bottomed flask, and approx. 200 mL triethyl orthoformate was added. The reaction was heated to 100° C. for 16 h, after which time the product had crashed out of solution as a white powder. The reaction cooled to room temperature, was filtered, and the solid washed with diethyl ether followed by pentane to give the amidinium salt (±)10 in 65% yield as a light fluffy white powder without further purification. Crystals suitable for X-ray analysis were achieved by vapor diffusion of n-pentane onto a CHCl$_3$ solution of (±)10. 1 H NMR (CDCl$_3$, 297K, 300 MHz): δ 0.97 (d, 2H, J=13.0 Hz), δ 1.14 (d, 2H, J=13.0 Hz), δ 1.47-2.09 (m, 22H), δ 2.66 (s, 2H), δ 4.77 (s, 2H), δ 7.43 (m, 2H), δ 7.51 (m, 6H), δ 8.84 (s, 1H). 13 C NMR (CDCl$_3$, 297K, 300 MHz): δ 26.39, δ 26.74, δ 30.00, δ 30.11, δ 30.36, δ 31.3, δ 31.7, δ 31.8, δ 32.2, δ 37.4, δ 37.8, δ 37.9, δ 56.7, δ 111.1, δ 116.3, δ 124.0, δ 129.1, δ 130.9, δ 145.3. HRMS (ESI-EMM): m/z=453.3247 ([M+BF$_4$]$^+$), Δ=3.0 ppm.

Synthesis of 2,2'-Bis(pivaloylamino)-1,1'-biphenyl, Compound 9Piv: 6.50 g (35.3 mmol) was weighed into an oven-dried 100 mL round-bottomed flask equipped with a stir bar and a septum-capped condenser, followed by N$_2$ purging. The system was charged with 70 mL dry THF and TEA (15.2 mL, 109 mmol). Pivaloyl chloride (11.3 mL, 91.7 mmol) was added dropwise via syringe to the above stirred solution, yielding a white precipitate. System was heated to reflux (80° C.) for 3 h under a N$_2$ atmosphere. The precipitate was filtered and washed with THF. Filtrate and washes were combined and liquids removed on a rotary evaporator giving pure 9Piv as a white powder in 100% yield. 1 H NMR (CDCl$_3$, 297K, 300 MHz): δ 1.00 (s, 18H), δ 7.18 (s, 2H), δ 7.24 (m, 4H), δ 7.46-8.11.
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(m, 2H), δ 8.32 (d, 2H, J=8.2 Hz). 13C NMR (CDCl3, 297K, 250 MHz): δ 27.33, δ 39.80, δ 122.15, δ 124.79, δ 128.36, δ 129.76, δ 129.94, δ 136.25, δ 176.90.

Synthesis of (R)-2,2'-bis(pivaloylamino)-1,1'-binaphthyl, Compound 9np. An oven-dried 500 mL round-bottom flask equipped with a stir bar was charged with LAH (8.03 g, 211 mmol) in a dry box. In a 100 mL oven-dried flask, 9np (12.44 g, 35.3 mmol) was dissolved in 200 mL dry THF under N2, which was then slowly cannula transferred to the stirred LAH suspension. To the reaction vessel was attached an oven-dried septum-capped condenser with a N2 inlet. The suspension was heated to reflux (80° C.) for 3.5 days, followed by careful quenching with H2O until fizzing ceased to be visible. 2M NaOH(aq) was added until a clear THF layer could be seen. The slurry was filtered and the solid washed with THF. The layers were separated and the aqueous layer was washed twice with 200 mL THF. Organic layers were combined, dried over MgSO4, and solvent removed to yield a mixture primarily consisting of starting material and product. The product was isolated as a clear oil in 91% yield by column chromatography (SiO2, 7.5% EtOAc in hexanes). 1H NMR (CDCl3, 300 MHz): δ 0.82 (s, 18H), δ 2.84 (m, 4H), δ 3.68 (t, 2H, J=5.1 Hz), δ 6.73 (m, 4H), δ 7.00-7.03 (m, 2H), δ 7.11-7.17 (m, 4H), δ 7.25 (d, J=9.3 Hz, 2H), δ 7.75 (m, 2H), δ 7.85 (d, J=9.0 Hz). 13C NMR (CDCl3, 297K, 250 MHz): δ 27.8, δ 32.1, δ 55.8, δ 110.2, δ 116.6, δ 123.7, δ 129.2, δ 130.8, δ 146.8. Synthesis of (R)-2,2'-bis(neopentylamino)-1,1'-biphenyl, Compound (R)-12Piv. See synthesis of (R)-12Av.

9Av (12.44 g, 35.3 mmol) was dissolved in 200 mL dry THF under N2, which was then slowly cannula transferred to the stirred LAH suspension. To the reaction vessel was attached an oven-dried septum-capped condenser with a N2 inlet. The suspension was heated to reflux (80° C.) for 3.5 days, followed by careful quenching with H2O until fizzing ceased to be visible. 2M NaOH(aq) was added until a clear THF layer could be seen. The slurry was filtered and the solid washed with THF. The layers were separated and the aqueous layer was washed twice with 200 mL THF. Organic layers were combined, dried over MgSO4, and solvent removed to yield a mixture primarily consisting of starting material and product. The product was isolated as a clear oil in 91% yield by column chromatography (SiO2, 7.5% EtOAc in hexanes). 1H NMR (CDCl3, 300 MHz): δ 0.77 (s, 18H), δ 4.00 (s, 4H), δ 7.38 (m, 2H), δ 7.51 (m, 6H), δ 8.83 (s, 1H). 13C NMR (CDCl3, 297K, 300 MHz): δ 27.1, δ 34.0, δ 66.4, δ 122.9, δ 130.0, δ 130.4, δ 130.6, δ 132.8, δ 146.3, δ 171.3. ESI-MS (m/z): calculated 335.2487, measured 335.2495.

Synthesis of neopentyl-substituted amidinium salt (±)-11. 9np (10.3 g, 32.0 mmol) and NH4BF4 (3.36 g, 32.1 mmol) were combined in a 500 mL round-bottom flask equipped with a stir bar and topped with a condenser. 20 mL of triethyl orthoformate was added, and the system was heated to ca. 100° C. for 20 h. After cooling to room temperature, pentane was added (ca. 100 mL). The suspension was filtered, and the crystals washed with pentane to give (±)-11 as a white powdery solid in 93% yield. 1H NMR (CDCl3, 300 MHz): δ 0.77 (s, 18H), δ 4.00 (s, 4H), δ 7.38 (m, 2H), δ 7.51 (m, 6H), δ 8.83 (s, 1H). 13C NMR (CDCl3, 297K, 300 MHz): δ 27.1, δ 34.0, δ 66.4, δ 122.9, δ 130.0, δ 130.4, δ 130.6, δ 132.8, δ 146.3, δ 171.3. ESI-MS (m/z): calculated 335.2487, measured 335.2495.

(R)-12Np. See synthesis of (±)-11. The product did not crash out of reaction, but was purified by removing excess orthoester under vacuum, dissolving in minimal CH2Cl2, and crashing out with pentane to give (R)-13 in 65%
yield as a light fluffy off-white powder. 1 H NMR (CDCl3, 297K, 300 MHz): δ 0.64 (s, 18H), δ 4.01 (d, J=13.8 Hz, 2H), δ 4.20 (d, J=13.8 Hz, 2H), δ 7.07 (d, J=8.4 Hz, 2H), δ 7.34 (t, J=7.2 Hz, 2H), δ 7.62 (m, 4H), δ 8.00 (d, J=8.1 Hz, 2H), δ 8.12 (d, J=9.0 Hz), δ 8.85 (s, lH).

13 C NMR (CDCl3, 297K, 300 MHz): δ 27.0, δ 33.9, δ 65.8, δ 120.2, δ 124.2, δ 126.4, δ 127.8, δ 128.1, δ 128.8, δ 131.5, δ 132.9, δ 147.5, δ 173.1. ESI-MS (m/z): calculated 435.2800, measured 435.2794.

Synthesis of [NHC-Pd(Cl)2]+, Compound 15. Compound 15 was synthesized according to an adaptation of a literature procedure, see Jensen & Sigman (2003) Org. Lett. 5:63-65. A 50 mL round-bottomed flask was charged with 100 mg (0.16 mmol) (±)14 and 2.0 mL 2.0 M ethereal HCI. The color instantly changed to bright yellow-orange. 8.0 mL ether was added and the resultant suspension stirred for 1 h. Volatiles were removed in vacuo leaving pure 15 as a bright yellow-orange powder in quantitative yield. Compound 15 could be recrystallized by taking up in a small amount of toluene and crashing out with excess n-pentane (87%). Crystals suitable for X-ray analysis were achieved by vapor diffusion of n-pentane onto a CH2Cl2 solution of 15. 1 H NMR (CDCl3, 297K, 500 MHz): δ 0.67 (d, J=12.5 Hz, 2H), δ 0.94 (d, J=12.5 Hz, 2H), δ 1.26-2.37 (m, 20H), δ 4.5 (br. s, lH), δ 4.8-5.7 (br. n, 3H), δ 7.27 (br. s, 2H), δ 7.34 (t, J=7.5 Hz, 4H), δ 7.45 (dd, J=7.5, 1.5 Hz, 2H). 13 C NMR (CDCl3, 297K, 500 MHz): δ 26.6, δ 26.98, δ 30.13, δ 30.49, δ 31.25, δ 35.81, δ 36.86, δ 37.24 (br.), δ 37.52, δ 68.10 (br.), δ 127.15, δ 127.28, δ 127.42, δ 127.79, δ 135.47, δ 145.78, δ 201.41. MS (ESI): m/z=(highest intensity peaks listed): 609.1 ([M-CW), 650.2 ([M+Li]+), 667.2 ([M+Na]+), 1255.3 ([M2-CW).
cooling to −78°C, 15 mL dry THF was added, and the solution was stirred overnight as the temperature slowly warmed to room temperature. Solvent was removed in a dry box, and residue was taken up in CH₂Cl₂. Filtration through a plug of celite gave a clear yellow/brown solution. Addition of AgOAc and stirring for 24 h gave a brown suspension with white ppt present. Filtration through celite, removal of solvent in vacuo, taking up in minimal CH₂Cl₂, and addition of excess hexanes gave a semi-cloudy light brown solution. Removal of small amount of solvent in vacuo promoted precipitation of an off-white solid, which was collected by filtration as pure (±)18 by NMR as a light tan powder in 18% yield. ³¹H NMR (CDCl₃, 297K, 300 MHz): δ 0.58 (s, 18H), δ 2.03 (s, 3H), δ 3.90 (d, 2H, J=13.4 Hz), δ 4.20 (d, 2H, J=13.4 Hz), δ 7.13 (m, 2H), δ 7.30 (m, 6H). ¹³C NMR (CDCl₃, 297K, 500 MHz): δ 27.8, δ 33.9, δ 71.9, δ 123.3, δ 127.8, δ 128.8, δ 129.0, δ 135.6, δ 149.5, δ 179.3 (s, carbene). MS (MALDI-TOF) (m/z): 335.3 ((NHC₆H⁺), 441.0 (NHC₄Ag⁺), 774.9 ((NHC₄)$_{2}$Ag⁺).

Crystallographic Experimental Data for Compound (±) 10:
A colorless crystal with approximate dimensions 0.20 x 0.13 x 0.03 mm³ was selected under oil under ambient conditions and attached to the tip of a nylon loop. The crystal was mounted in a stream of cold nitrogen at 100 K and centered in the X-ray beam by using a video camera.

The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα (λ=0.71073 Å) radiation and the diffractometer to crystal distance of 4.9 cm.

The initial cell constants were obtained from three series of ω scans at different starting angles. Each series consisted of 30 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 20 seconds per frame. A total of 59 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 1,469 strong reflections from the actual data collection.

The data were collected by using the hemisphere data collection routine. The reciprocal space was surveyed to the extent of a full sphere to a resolution of 0.80 Å. A total of 22,836 data were harvested by collecting three sets of frames with 0.3° scans in ω with an exposure time 30 sec per frame. These highly redundant datasets were corrected for Lorentz and polarization effects. The absorption correction was based on fitting a function to the empirical transmission surface as sampled by multiple equivalent measurements, as noted earlier.

Structure Solution and Refinement: The systematic absences in the diffraction data were uniquely consistent for the space group P2₁/c that yielded chemically reasonable and computationally stable results of refinement.

A successful solution by the direct methods provided most non-hydrogen atoms from the E-map. The remaining non-hydrogen atoms were located in an alternating series of least-squares cycles and difference Fourier maps. All non-hydrogen atoms were refined with anisotropic displacement coefficients unless otherwise indicated. All hydrogen atoms were included in the structure factor calculation at idealized positions and were allowed to ride on the neighboring atoms with relative isotropic displacement coefficients.

The allyl group is disordered over two positions in a 73:27 ratio and was refined with restraints. There is also one-half of solvated hexane molecule per Pd complex in the lattice. The solvent molecule was refined isotropically with restraints.

The final least-squares refinement of 372 parameters against 2,791 data resulted in residuals R (based on F² for 1≤2r) and wR (based on F² for all data) of 0.0495 and 0.1315, respectively. The final difference Fourier map was featureless. See FIG. 1 for the resulting molecular structure with thermal ellipsoids shown at 50% probability.

Crystallographic Experimental Data for Compound 14: A yellow crystal with approximate dimensions 0.46 x 0.25 x 0.15 mm³ was selected under oil under ambient conditions and attached to the tip of a nylon loop. The crystal was mounted in a stream of cold nitrogen at 200(2) K and centered in the X-ray beam by using a video camera.

The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα (λ=0.71073 Å) radiation and the diffractometer to crystal distance of 4.9 cm.

The initial cell constants were obtained from three series of ω scans at different starting angles. Each series consisted of 20 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 10 seconds per frame. A total of 67 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 5,499 strong reflections from the actual data collection.

The data were collected by using the hemisphere data collection routine. The reciprocal space was surveyed to the extent of a full sphere to a resolution of 0.80 Å. A total of 12,112 data were harvested by collecting four sets of frames with 0.25° scans in ω with an exposure time 38 sec per frame. These highly redundant datasets were corrected for Lorentz and polarization effects. The absorption correction was based on fitting a function to the empirical transmission surface as sampled by multiple equivalent measurements, as noted earlier.

The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα (λ=0.71073 Å) radiation and the diffractometer to crystal distance of 4.9 cm.
The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα (λ=0.71073 Å) radiation and the diffractometer to crystal distance of 4.9 cm.

The initial cell constants were obtained from three series of scans at different starting angles. Each series consisted of 20 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 10 seconds per frame. A total of 54 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 11,665 strong reflections from the actual data collection.

The reflections were harvested by collecting three sets of frames with 0.3° scans at different starting angles. Each series consisted of 20 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 10 seconds per frame. A total of 54 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 11,665 strong reflections from the actual data collection.

The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα radiation and the diffractometer to crystal distance of 4.9 cm.

The initial cell constants were obtained from three series of scans at different starting angles. Each series consisted of 20 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 10 seconds per frame. A total of 54 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 11,665 strong reflections from the actual data collection.

The reflections were harvested by collecting three sets of frames with 0.3° scans in ω with an exposure time 36 sec per frame. These highly redundant datasets were corrected for Lorentz and polarization effects. The absorption correction was based on fitting a function to the empirical transmission surface as sampled by multiple equivalent measurements.

Structure Solution and Refinement: The systematic absences in the diffraction data were consistent for the space groups Cc and C2/c. The E-statistics strongly suggested the centrosymmetric space group C2/c that yielded chemically reasonable and computationally stable results of refinement.

A successful solution by the direct methods provided most non-hydrogen atoms from the E-map. The remaining non-hydrogen atoms were located in an alternating series of least-squares cycles and difference Fourier maps. All non-hydrogen atoms were refined with anisotropic displacement coefficients. All hydrogen atoms were included in the structure factor calculation at idealized positions and were allowed to ride on the neighboring atoms with relative isotropic displacement coefficients.

There were two solvate molecules of dichloromethane and/or ether present in the asymmetric unit. A significant amount of time was invested in identifying and refining the disordered molecules. Bond length restraints were applied to model the molecules but the resulting isotropic displacement coefficients suggested the molecules were mobile. In addition, the refinement was computationally unstable. The "SQUEEZE" option in the computer program PLATON was used to correct the diffraction data for diffuse scattering effects and to identify the solvate molecule. PLATON is a crystallographic modeling tool, copyright 1980-2005 by A. L. Spek, and can be obtained for free from Utrecht University, Padualaan 8, 3584 CH, Utrecht, The Netherlands. PLATON calculated the upper limit of volume that could be occupied by the solvent to be 1558 Å³, or 21% of the unit cell volume. The program calculated 689 electrons in the unit cell for the diffuse species. This approximately corresponds to two molecules of dichloromethane per Pd complex in the asymmetric unit (672 electrons). It is very likely that this solvate molecules are disordered over several positions. Note that all derived results in the following tables are based on the known contents. No data are given for the diffusely scattering species.

The final least-squares refinement of 414 parameters against 7613 data resulted in residuals R (based on F² for I≥2σ(I)) and wR (based on F² for all data) of 0.0346 and 0.0960, respectively. The resulting molecular structure is depicted in Fig. 4. The probability ellipsoids are at 30%. The content of the asymmetric unit of 12 is depicted in Fig. 5. The hydrogen atoms have been omitted in both Figs. 4 and 5 for clarity.

Crystallographic Experimental Data for Compound 16: A yellow crystal with approximate dimensions 0.36x0.25x0.18 mm² was selected under oil under ambient conditions and attached to the tip of a nylon loop. The crystal was mounted in a stream of cold nitrogen at 100 K and centered in the X-ray beam by using a video camera.

The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα radiation and the diffractometer to crystal distance of 4.9 cm.

The initial cell constants were obtained from three series of scans at different starting angles. Each series consisted of 20 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 10 seconds per frame. A total of 54 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 11,665 strong reflections from the actual data collection.

The reflections were harvested by collecting three sets of frames with 0.3° scans in ω with an exposure time 36 sec per frame. These highly redundant datasets were corrected for Lorentz and polarization effects. The absorption correction was based on fitting a function to the empirical transmission surface as sampled by multiple equivalent measurements.
The data were collected by using the full sphere data collection routine to survey the reciprocal space to the extent of a full sphere to a resolution of 0.80 Å. A total of 27822 data were harvested by collecting three sets of frames with 0.3° scans in ω with an exposure time 36 sec per frame. These highly redundant datasets were corrected for Lorentz and polarization effects. The absorption correction was based on fitting a function to the empirical transmission surface as sampled by multiple equivalent measurements.

Structure Solution and Refinement: The systematic absences in the diffraction data were uniquely consistent for the space group P2₁/n that yielded chemically reasonable and computationally stable results of refinement.

A successful solution by the direct methods provided most non-hydrogen atoms from the E-map. The remaining non-hydrogen atoms were located in an alternating series of least-squares cycles and difference Fourier maps. All non-hydrogen atoms were refined with anisotropic displacement coefficients unless specified otherwise. All hydrogen atoms were included in the structure factor calculation at idealized positions and were allowed to ride on the neighboring atoms with relative isotropic displacement coefficients.

The final least-squares refinement of 344 parameters against 6950 data resulted in residuals R (based on F² for all data) of 0.1034 and 0.2651, respectively.

Crystallographic Experimental Data for Compound 18: A colorless crystal with approximate dimensions 0.299 x 0.23 x 0.19 mm³ was selected under oil under ambient conditions and attached to the tip of a nylon loop. The crystal was mounted in a stream of cold nitrogen at 200(2) K and centered in the X-ray beam by using a video camera.

The crystal evaluation and data collection were performed on a Bruker CCD-1000 diffractometer with Mo Kα (λ=0.71073 Å) radiation and the diffractometer to crystal distance of 4.9 cm.

The initial cell constants were obtained from three series of ω scans at different starting angles. Each series consisted of 20 frames collected at intervals of 0.3° in a 6° range about ω with the exposure time of 15 seconds per frame. A total of 206 reflections was obtained. The reflections were successfully indexed by an automated indexing routine built in the SMART program. The final cell constants were calculated from a set of 10935 strong reflections from the actual data collection.

The data were collected by using the hemisphere data collection routine. The reciprocal space was surveyed to the extent of a full sphere to a resolution of 0.80 Å. A total of 24283 data were harvested by collecting three sets of frames with 0.25° scans in ω with an exposure time 25 sec per frame. These highly redundant datasets were corrected for Lorentz and polarization effects. The absorption correction was based on fitting a function to the empirical transmission surface as sampled by multiple equivalent measurements.

Structure Solution and Refinement: The systematic absences in the diffraction data were uniquely consistent for the space group P2₁/n that yielded chemically reasonable and computationally stable results of refinement.

What is claimed is:

1. N-heterocyclic carbenes, salts thereof, and metal complexes thereof, comprising formula I:

wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted mono- or polycyclic cycloalkenyl, cycloalkyl, heterocycloalkyl, and heteroarylalkenyl having up to 34 atoms within each of ring A or ring B, provided that A and B are not phenyl, naphthalenyl, or 9,10-dihydrophenanthrenyl; wherein heteroatoms in either of ring A or ring B, if any, are independently selected from the group consisting of N, O, S, and P; R₁, R₂, R₃, R₄, R₅, and R₆ are independently selected from the group consisting of hydrogen; C₁ to C₉₅ substituted or unsubstituted, linear or branched alkyl, alkynyl, and aralkyl, aryl and heteroaromatic; or R₁, R₂, and R₃, including or excluding the carbon to which they are attached and independent of R₄, R₅, and R₆, R₄, R₅, and R₆, including or excluding the carbon to which they are attached and independent of R₁, R₂, R₃, and R₄, define a C₃ to C₉₅ substituted or unsubstituted, mono- or polycyclic cycloalkyl, cycloalkenyl, cycloalkynyl, cycloalkyl, heteroaryl, heterocycloalkenyl, heteroarylalkenyl, wherein heteroatoms in any any, are independently selected from the group consisting of N, O, S, and P; wherein substituents on substituted moieties are selected from the group consisting of halogen; linear or branched C₁-C₉₅-alkyl, alkynyl, or aralkyl; C₃-C₅₂-cycloalkyl, cycloalkenyl, or cycloalkynyl.
The carbenes of claim 1, comprising an amidinium salt.

7. The carbenes of claim 1, comprising formula III:

wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted mono- or poly cyclic cycloalkyl, cycloalkyl, heterocycloalkyl, and heterocycloalkenyl having up to 34 atoms within each of ring A or ring B; wherein heteroatoms in either of ring A or ring B, if any, are independently selected from the group consisting of N, O, S, and P; R₁, R₂, R₃, R₄, R₅, and R₆ are independently selected from the group consisting of hydrogen; C₁ to C₆₀ substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl; or R₁, R₂, R₃, R₄, R₅, and R₆, including or excluding the carbon to which they are attached and independent of R₁, R₂, R₃, R₄, R₅, and R₆, including or excluding the carbon to which they are attached and independent of R₁, R₂, R₃, R₄, R₅, and R₆, define a C₃ to C₆₀ substituted or unsubstituted, mono- or poly cyclic cycloalkyl, cycloalkenyl, cycloalkynyl, cycloaryl, heterocycloalkyl, heterocycloalkenyl, heterocycloalkynyl, wherein heteroatoms in if any, are independently selected from the group consisting of N, O, S, and P; wherein substituents on substituted moieties are selected from the group consisting of halogen; linear or branched C₁₋C₁₂-alkyl, alkenyl, and alkynyl; C₅₋C₁₂-cycloalkyl, cycloalkenyl, or cycloalkynyl; mono- or poly cyclic aryl, and mono- or poly cyclic heteroaryl having up to 5 heteroatoms selected from N, O, S, and P; M is hydrogen or a metal; L when present is a ligand coordinated to M; X is a counter-anion; and “a” is an integer.

8. The carbenes of claim 7, wherein M is a main group metal.

9. The carbenes of claim 7, wherein M is a transition group metal.

10. The carbenes of claim 7, wherein M is hydrogen or a metal selected from the group consisting of Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Os, Ir, Pt, and Au.

11. The carbenes of claim 7, wherein X is selected from the group consisting of tetrafluoroborate, hexafluorophosphate, and halide.

12. N-heterocyclic carbenes comprising formula III:

wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted mono- or poly cyclic cycloalkenyl, cycloalkyl, heterocycloalkyl, and heterocycloalkenyl having up to 34 atoms within each of ring A or ring B; wherein heteroatoms in either of ring A or ring B, if any, are independently selected from the group consisting of N, O, S, and P; R₁, R₂, R₃, R₄, R₅, and R₆ are independently selected from the group consisting of hydrogen; C₁ to C₆₀ substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl; or R₁, R₂, R₃, R₄, R₅, and R₆, including or excluding the carbon to which they are attached and independent of R₁, R₂, R₃, R₄, R₅, and R₆, define a C₃ to C₆₀ substituted or unsubstituted, mono- or poly cyclic cycloalkyl, cycloalkenyl, cycloalkynyl, cycloaryl, heterocycloalkyl, heterocycloalkenyl, heterocycloalkynyl, wherein heteroatoms in if any, are independently selected from the group consisting of N, O, S, and P; wherein substituents on substituted moieties are selected from the group consisting of halogen; linear or branched C₁₋C₁₂-alkyl, alkenyl, and alkynyl; C₅₋C₁₂-cycloalkyl, cycloalkenyl, or cycloalkynyl; mono- or poly cyclic aryl, and mono- or poly cyclic heteroaryl having up to 5 heteroatoms selected from N, O, S, and P; M is hydrogen or a metal selected from the group consisting of Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Os, Ir, Pt, and Au; L when present is a ligand coordinated to M; X is a counter-anion; and “a” is an integer.

13. N-heterocyclic carbenes, salts thereof, and metal complexes thereof, comprising formula I:
wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted pyridine, 9H-fluoren-9-one, 5-phenylthiazole, 1,2,3,4-tetrahydronaphthalene, phenanthrene, thiophene, anthracene-9,10-dione, 9-(tert-butyl)-7-methyl-6H-benzo[c]chromene, 5-methyl-pyrazole, 10-hydroxy-3,4-dihydro-1H-benzo[g]isochromene-1,6,9-trione, and thiazole; R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen; C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl; or R1, R2, and R3, including or excluding the carbon to which they are attached and independent of R4, R5, and R6, including or excluding the carbon to which they are attached and independent of R1, R2, and R3, define a C3 to C60 substituted or unsubstituted, mono- or polycyclic cycloalkyl, cycloalkenyl, cycloalkynyl, cycloaryl, heterocycloalkyl, heterocycloalkenyl, heterocycloalkynyl, wherein heteroatoms in if any, are independently selected from the group consisting of N, O, S, and P; wherein substituents on substituted moieties are selected from the group consisting of halogen; linear or branched C1-C12 cycloalkyl, cycloalkenyl, cycloalkynyl, cycloaryl, heterocycloalkyl, heterocycloalkenyl, heterocycloalkynyl, monocyclic cycloalkyl, cycloalkenyl, cycloalkynyl, heterocycloalkyl, and mono- or polycyclic heteroaryl having up to 5 heteroatoms selected from N, O, S, and P; and salts thereof; and metal complexes thereof.

14. The carbenes of claim 13, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen; C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl.

15. The carbenes of claim 13, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen and C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl.

16. The carbenes of claim 13, wherein R1, R2, R3, R4, and R5 are independently selected from the group consisting of hydrogen and the carbon to which they are attached, and R4, R5, and R6, the carbon to which they are attached, are adamantyl.

17. The carbenes of claim 13, comprising an amidinium salt.

18. The carbenes of claim 13, wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted pyridine, phenanthrene, thiophene, and thiazole.

19. The carbenes of claim 18, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen; C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl.

20. The carbenes of claim 18, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen and C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, and the carbon to which they are attached, and R4, R5, and R6, the carbon to which they are attached, are adamantyl.

21. The carbenes of claim 18, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen; C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl.

22. The carbenes of claim 18, comprising an amidinium salt.

23. The carbenes of claim 13, wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted pyridine, 9H-fluoren-9-one, 5-phenylthiazole, 1,2,3,4-tetrahydronaphthalene, phenanthrene, thiophene, anthracene-9,10-dione, 9-(tert-butyl)-7-methyl-6H-benzo[c]chromene, 5-methyl-pyrazole, 10-hydroxy-3,4-dihydro-1H-benzo[g]isochromene-1,6,9-trione, and thiazole; R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen; C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl, aryl and heteroaryl.

24. The carbenes of claim 23, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen; C1 to C60 substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl.

25. The carbenes of claim 23, wherein R1, R2, R3, R4, R5, and R6 are independently selected from the group consisting of hydrogen and the carbon to which they are attached, and R4, R5, and R6, the carbon to which they are attached, are adamantyl.

26. The carbenes of claim 23, wherein R1, R2, R3, and the carbon to which they are attached, and R4, R5, and R6, the carbon to which they are attached, are adamantyl.

27. The carbenes of claim 23, comprising an amidinium salt.

28. The carbenes of claim 23, comprising formula III:
monocyclic aryl, and mono- or polycyclic heteroaryl having up to 5 heteroatoms selected from N, O, S, and P; M is hydrogen or a metal; L when present is a ligand coordinated to M; X is a counter-anion; and “a” is an integer.

29. The carbenes of claim 28, wherein M is a main group metal.

30. The carbenes of claim 28, wherein M is a transition group metal.

31. The carbenes of claim 28, wherein M is hydrogen or a metal selected from the group consisting of Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Os, Ir, Pt, and Au.

32. The carbenes of claim 28, wherein X is selected from the group consisting of tetrafluoroborate, hexafluorophosphate, and halide.

33. N-heterocyclic carbenes comprising formula III:

\[
\begin{array}{c}
\text{Ring A} \\
\begin{array}{c}
\text{N} \\
\text{M} \\
\text{L}_{0.5}
\end{array}
\end{array}
\begin{array}{c}
\text{Ring B} \\
\begin{array}{c}
\text{N} \\
\text{R}^1 \text{R}^2 \text{R}^3 \text{R}^4 \text{R}^5
\end{array}
\end{array}
\]

wherein ring A and ring B are independently selected from the group consisting of substituted or unsubstituted pyridine, 9H-fluoren-9-one, 1,2,3,4-tetrahydronaphthalene, phenanthrene, thiophene, anthracen-9,10-dione, 9-(tert-butyl)-7-methyl-6H-benzo[e]chromene, 5-methyl-pyrrole, 10-hydroxy-3,4-dihydro-1H-benzo[g]isochromene-1,6,9-trione, and thiiazole; 

R\(^1\), R\(^2\), R\(^3\), R\(^4\), and R\(^5\) are independently selected from the group consisting of hydrogen; C\(_2\) to C\(_{60}\) substituted or unsubstituted, linear or branched alkyl, alkenyl, and alkynyl and heteroaryl, aryl; or R\(^1\), R\(^2\), and R\(^3\), including or excluding the carbon to which they are attached and independent of R\(^4\), R\(^5\), and R\(^6\), including or excluding the carbon to which they are attached and independent of R\(^1\), R\(^2\), and R\(^3\), define a C\(_3\) to C\(_{60}\) substituted or unsubstituted, mono- or bicyclic cycloalkyl, cycloalkenyl, cycloalkynyl, cycloaryl, heterocycloalkyl, heterocycloalkenyl, heterocycloalkynyl, wherein heteroatoms in any position are independently selected from the group consisting of N, O, S, and P; wherein substituents on substituted moieties are selected from the group consisting of halogen; linear or branched C\(_1\)-C\(_{12}\)-alkyl, alkenyl, or alkynyl; C\(_5\)-C\(_{10}\)-cycloalkyl, cyanoalkyl, or cyanoalkynyl; mono- or bicyclic aryl, and mono- or bicyclic heteroaryl having up to 5 heteroatoms selected from N, O, S, and P; M is hydrogen or a metal selected from the group consisting of Fe, Co, Ni, Cu, Ru, Rh, Pd, Ag, Os, Ir, Pt, and Au; L when present is a ligand coordinated to M; X is a counter-anion; and “a” is an integer.